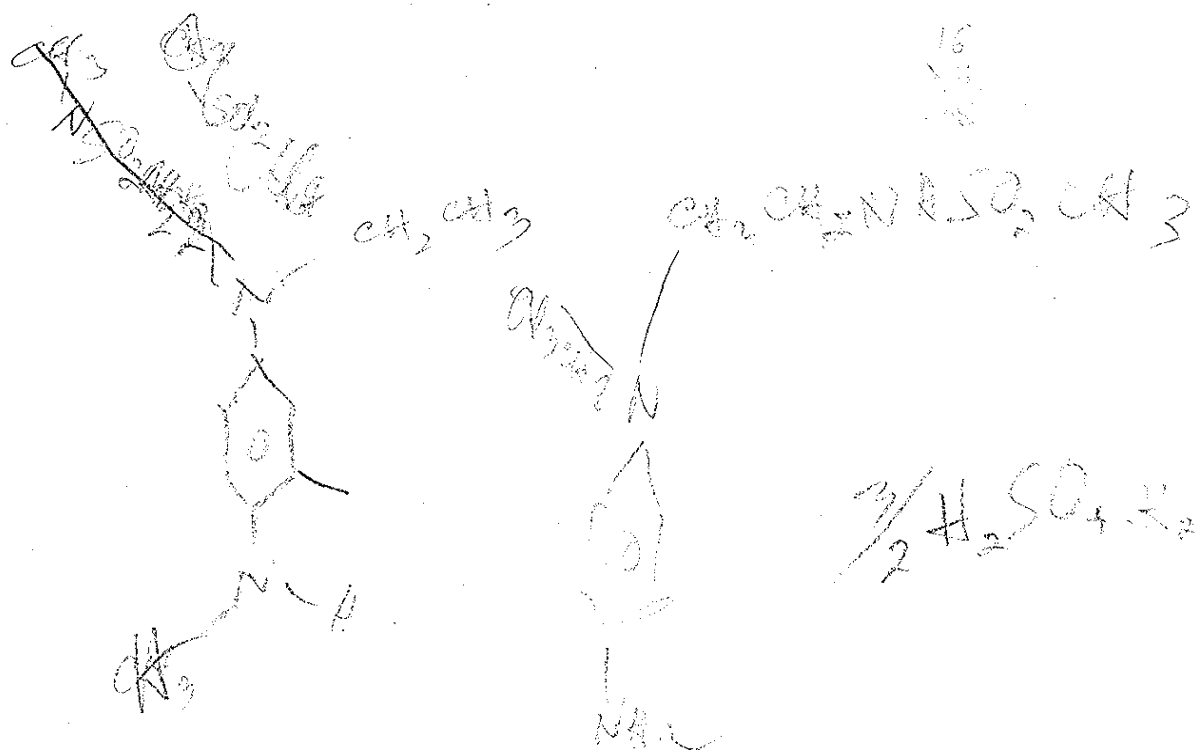


$$NH_4CO_3 = 14 + 14 + 3 \times 16 = 106 \text{ g/mol}$$



$$\frac{3}{2} \text{H}_2\text{SO}_4 + \text{K}_2\text{O}$$

12
11C
1940
19

$$\begin{array}{r} 16 \\ 02 \\ \hline 18 \end{array} \quad \begin{array}{r} 11 \\ 09 \\ \hline 20 \end{array} \quad \begin{array}{r} 91 \\ 10 \\ \hline 101 \end{array}$$

12

257
147

$$2 + 42 + 2^2 = 49$$

4099/116C

1000

$$\frac{50g}{100g} \times \frac{mol}{401g} = 0.12336 \frac{mol}{g}$$

FIG. 1

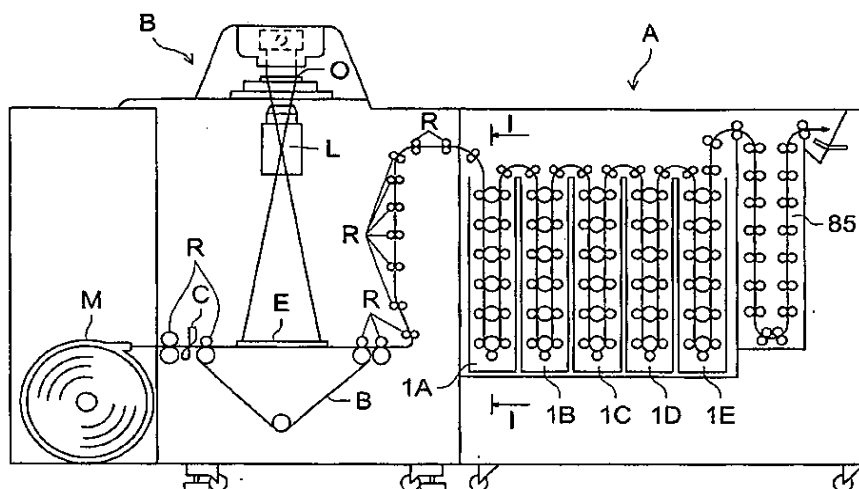


FIG. 1

